



Consider a first order chemical reaction in which Chemical 1 is converted to Chemical 2

 $C_1 \rightarrow C_2$

Rate =
$$k C_1$$

$$\frac{dC_1}{dt} = -kC_1$$

 $C_1(t) = C_1(0)e^{-kt}$

 $C_2(t) = C_1(0) - C_1(t)$

 $C_2(t) = C_1(0) (1-e^{-kt})$

$M_2(t) = M_1(0) (1-e^{-kt})$

For values of chemical rxn time constant $\beta = kt < 0.01$, $M_2 \approx M_1(0)kt$

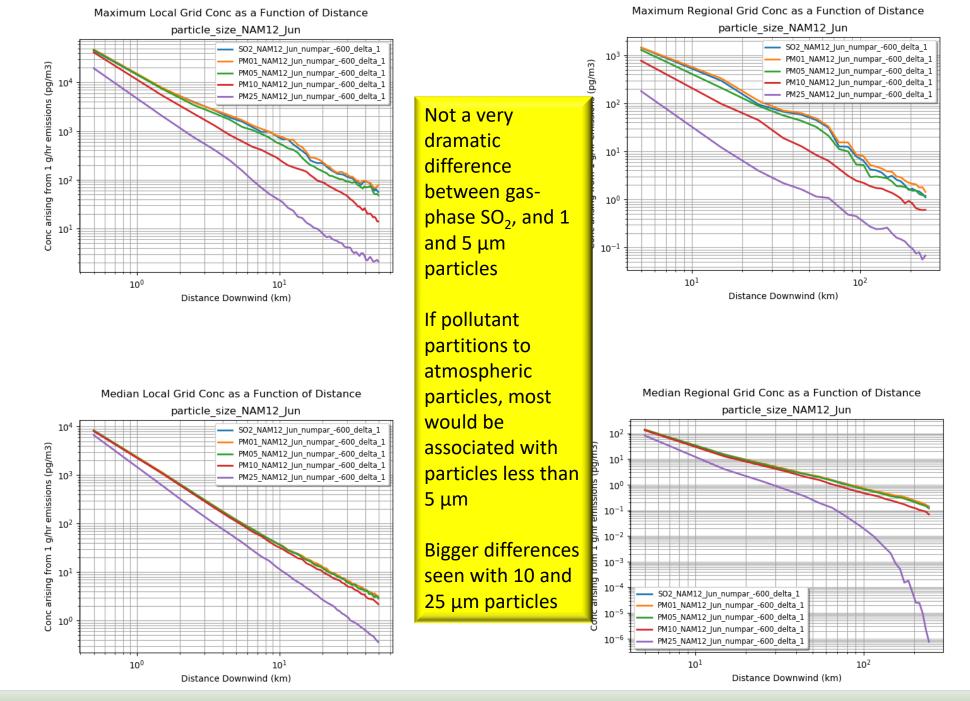
If the molecular weight (W_2) of C_2 is different from the molecular weight (W_1) of C_1 , then the mass of C_2 is increased by the ratio of molecular weights, i.e.,

 $M_2(t) = M_1(0) (1-e^{-kt}) (W_2/W_1)$

For example: $SO_2 \rightarrow SO_4^{-2}$

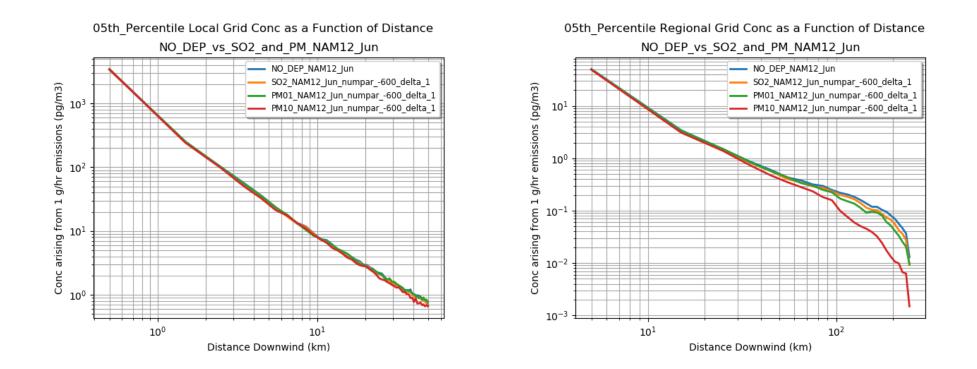
Molecular weight of $SO_2 = 64$ Molecular weight of $SO_4^{-2} = 96$ Ratio = 1.5

For every amount of mass that SO_2 is depleted by the reaction, the mass of SO_4^{-2} increases by 1.5 x that amount of mass (the conversion adds two extra oxygen atoms to the molecule)



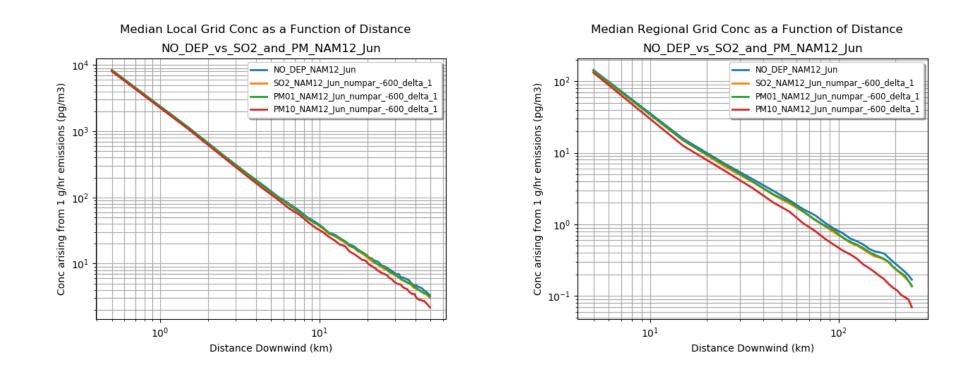
HYSPLIT Simulations for ALOHA Chemicals (12/04/2018)

Compare simulation with no deposition with simulation of SO2, PM01 and PM10 with default deposition parameters



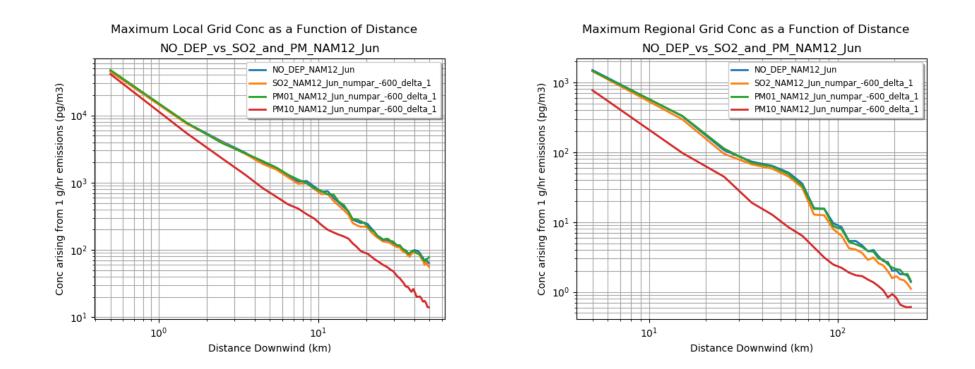
For 5th percentile, where one might expect to see the consequences of deposition (e.g., when it is raining), little difference except for large distances with large particles (10 μm)

Compare simulation with no deposition with simulation of SO2, PM01 and PM10 with default deposition parameters



For median concentrations, little difference except for large particles (10 μm)

Compare simulation with no deposition with simulation of SO2, PM01 and PM10 with default deposition parameters



For maximum concentrations, little difference except for large particles (10 μm)

Example of overall impact of wet/dry deposition (for SO₂)

Statistical Distribution of Hourly Concentration Values

local_grid_9.5_km

